

Name _____

Chemical Reactions Review

1. What is the law of conservation of mass?
2. Why do we need to balance equations?
3. What do each of these symbols mean:
 - a) aq
 - b) s
 - c) l
 - d) g

Remember that Hydrogen (H₂), Nitrogen (N₂), Oxygen (O₂), Fluorine (F₂), Chlorine (Cl₂), Bromine (Br₂), and Iodine (I₂) are diatomic.

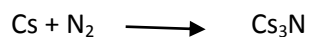
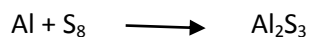
Write the reaction from these word equations and balance these equations as well.

Magnesium Sulfate plus Copper II Phosphate makes Magnesium Phosphate and Copper II Sulfate.

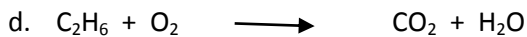
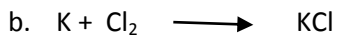
Calcium plus Sodium Nitrate makes Calcium Nitrate and Sodium.

What is a skeleton equation?

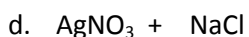
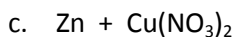
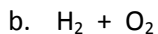
Balance these equations



4. What is a coefficient?
5. Why can't you change the subscript when balancing equations?
6. Know how to identify types of reactions and how to predict products.
7. Identify these types of reactions and balance them.



8. Identify the type of reaction and predict the products, balance the equation as well.



9. Know how to read an activity series

10. Here is the activity series

Mg

Al

Zn

Fe

Pb

Cu

Ag

Tell me if these reactions will occur:

- a. $\text{Zn} + \text{PbSO}_4$
- b. $\text{Cu} + \text{Mg}(\text{NO}_3)_2$
- c. $\text{Fe} + \text{AlPO}_4$
- d. $\text{Cu} + \text{AgNO}_3$
- e. $\text{Zn} + \text{AlPO}_4$