

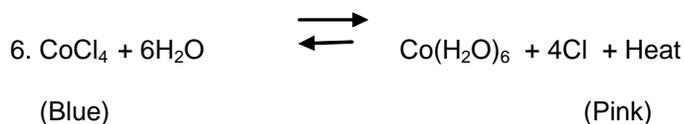
1. Equilibrium is a state of dynamic molecular behavior, meaning that
 - a. The amounts of product and reactant are equal
 - b. reactants continually turn into products at a progressively slower rate.
 - c. products continually turn into reactants at a progressively faster rate.
 - d. reactants turn into products and products turn into reactants at different rates.
 - e. reactants turn into products and products turn into reactants at equal rates.

2. A reversible reaction is:
 - a. a reaction that can occur in both forward and reverse directions
 - b. a reaction that alternates between going to the right and going to the left
 - c. a reaction that goes to completion
 - d. not very important to the study of equilibrium

3. A state in which the forward and reverse reactions balance each other because they take place at equal rates
 - a. reversible reaction
 - b. chemical equilibrium
 - c. solubility product constant
 - d. common ion effect
 - e. Le Chatalier's principle

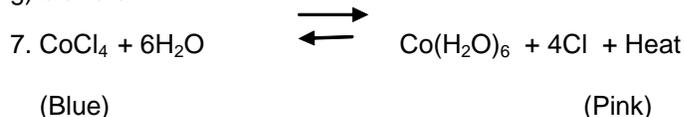
4. The best way to be certain that any chemical equilibrium will produce the largest amount of product is to:
 - a. increase the temp. and decrease the pressure at the same time
 - b. raise the temp.
 - c. increase the pressure
 - d. keep removing product

5. If stress is applied to a system at equilibrium, the system shifts in the direction that relieves the stress.
 - a. reversible reaction
 - b. Le Chatalier's principle
 - c. chemical equilibrium
 - d. law of chemical equilibrium
 - e. common ion effect



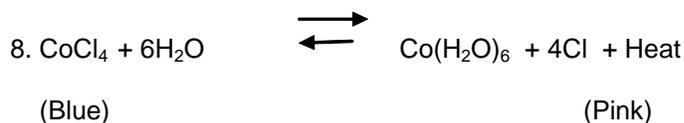
- If I add more Chloride ions this will cause the reaction to
- a) shift to the right (products)
 - b) Shift to the left (reactants)
 - c) stay the same
 - d) become blue

- e) become pink
- f) a and e
- g) b and d



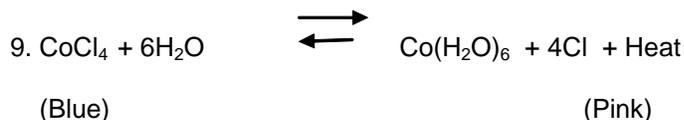
If I add more water this will cause the reaction to

- a) shift to the right (products)
- b) Shift to the left (reactants)
- c) stay the same
- d) become blue
- e) become pink
- f) a and e
- g) b and d



If I add heat to the reaction this will cause the reaction to

- a) shift to the right (products)
- b) Shift to the left (reactants)
- c) stay the same
- d) become blue
- e) become pink
- f) a and e
- g) b and d



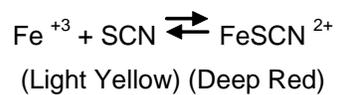
If I remove heat this will cause the reaction to

- a) shift to the right (products)
- b) Shift to the left (reactants)
- c) stay the same
- d) become blue
- e) become pink
- f) a and e
- g) b and d

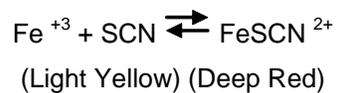
10. Increasing the concentration of NH_3 will shift the following reaction:



- a. to the left
- b. towards the products
- c. to the right
- d. towards reactants
- e. a and d
- g. b and c
- h. no effect



11. Adding AgNO_3 to the test tube removes SCN^- from the reaction, causing the reaction to
- produce more reactants.
 - produce a deeper red color.
 - A, and B are correct.
 - not enough information



12. The reaction you studied becomes a deeper red color when placed in an ice bath. This means that the forward reaction is
- There is not enough information given to answer the question.
 - endothermic.
 - exothermic.
 - None of the above are correct.