

Name \_\_\_\_\_

### **Spectral Lines**

Go to [chalmerschemistry.weebly.com](http://chalmerschemistry.weebly.com), scroll over to first semester and click on resources. Find the link that says line emission spectra explanation and click on it.

1. What are spectral lines?
2. What color lines are present for Hydrogen?
3. What color lines are present for Beryllium?
4. What was Bohr's radical idea about electrons?
5. What is quanta?
6. What idea from the Bohr model do we still keep?
7. What kind of orbits would have lower energy?
8. Do electrons actually move in circular orbits?
9. In order for an electron goes down an energy level what has to happen?
10. In order for an electron to go up an energy level what has to happen?
11. From the reading today how does line emission spectra show that atoms of the same element have different energy levels?